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Ion. Consists of a single atom with a
positive or negative charge resulting from
the loss of gain of one or more valence

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electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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To name a binary molecular compound, use the following guidelines: 1. Write the

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names of the elements in the order listed in the formula 2. Use prefixes appropriately to indicate the number of each kind of atom 3. End the name of the second element with the suffix -ide

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Answers Key First, confirm that the compound is a binary molecular compound. Name the elements in the order listed in the formula. Use prefixes to indicate the number of each kind of atom. Omit the prefix -mono when the formula contains only one atom of the first element in the name. The suffix of the name of the second element is -ide.

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1. when the name of the anion (X) ends in -ide the acid name begins with the prefix hydro-. The stem of the anion has the suffix -ic and is followed by the word acid (ex. HCl= hydrochloric acid) 2. when the

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anion name ends in -ite the acid name is
the stem of the anion with the suffix -ous
followed by the word acid (ex. H_2SO_3 --
X=sulfite = sulfurous acid)

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Writing the Names and Formulas for

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Bases. A base is a compound that produces OH^- in water. When naming a base, you name it like any other compound that starts with a regular or transition metal. Ex: NaOH = sodium hydroxide. When writing the formula for a base, remember to balance charges. Ex: magnesium hydroxide = $\text{Mg}(\text{OH})_2$

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Chapter 9 Chemical Names and
Formulas83 SECTION 9.3 NAMING
AND WRITING FORMULAS FOR
MOLECULAR COMPOUNDS (pages
268–270) This section explains the rules

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Answers and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

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~~ANSWERS KEY 9~~

SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC COMPOUNDS

I. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3. Name the following binary ionic compounds. sodium sulfide g. CuCl: h.

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snc.l. a. MnO₂ c. d. SrBr₂ 221 Chapter 9
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CHAPTER 9 Assessment 42. a. 2- b. 1+ c.

1- d. 3+ 43. a. 2+ b. 2+ c. 3+ d. 1+

44. a. barium ion b. iodide ion c. silver ion d.

mercury(II) ion 45. cyanide, CN^- and

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hydroxide, OH⁻ 46. a. hydroxide ion b.
lead(IV) ion c. sulfate ion d. oxide ion 47.
zero; A compound is electrically neutral.
48. The symbols for the cation and

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(20) Monatomic Ion. Consists of a single
atom with a positive or negative charge
resulting from the loss of gain of one or
more valence

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this chapter. A B; anion: any atom or group of atoms that has a negative charge:

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DIF: L1 REF: p. 253 OBJ: 9.1.1 Identify
the charges of monatomic ions by using
the periodic table, and name the ions.

STA: Ch.5.a 2. ANS: H PTS: 1 DIF: L1
REF: p. 254

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